

# Keywords: Autumn 2

## Science

### Topic Title: Quantitative Chemistry

#### Introduction to the topic:

**This topic will be very maths-based chemistry – remember your calculators! You will be calculating masses of reactants and products of chemical reactions, calculating concentrations and calculating moles (not the animal!).**

Keyword	Definition
<b>Avogadro's Constant</b>	<b>The number of atoms, molecules or ions in a mole of a given substance.</b>
<b>Concentration</b>	<b>The amount of substance (e.g. the mass) in a certain volume of a solution.</b>
<b>Conservation of Mass</b>	<b>The law of conservation of mass states that no atoms are lost or made during a chemical reaction so the mass of the products equals the mass of the reactants.</b>
<b>Relative Atomic Mass</b>	<b>The mean relative mass of the atoms of the different isotopes in an element. It is the number of times heavier an atom is than one-twelfth of a carbon-12 atom.</b>
<b>Mole</b>	<b>Chemical amounts are measured in moles. The mole is the unit for amount of substance. The symbol for the unit mole is mol.</b>
<b>Closed system</b>	<b>In chemistry, a system in which no substances can enter or leave during a reaction.</b>
<b>Relative Formula Mass</b>	<b>The sum of the relative atomic masses of the atoms in a chemical formula.</b>